



Standard Test Method for Sodium in Water by Atomic Absorption Spectrophotometry¹

This standard is issued under the fixed designation D4191; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope*

1.1 This test method covers the determination of low amounts of sodium in waters² having low solids content. The applicable range of this test method is from 0.20 to 3.0 mg/L when using the 589.6-nm resonance line. This range may be extended upward by dilution of an appropriate aliquot of sample or by use of the less-sensitive 330.2-nm resonance line (see Test Method D3561). Many workers have found that this test method is reliable for sodium levels to 0.005 mg/L, but use of this test method at this low level is dependent on the configuration of the aspirator and nebulizer system available in the atomic absorption spectrophotometer as well as the experience and skill of the analyst. The precision and bias data presented are insufficient to justify use of this test method in the 0.005 to 0.20-mg/L range.

1.2 This test method has been used successfully with spiked reagent water. It is the analyst's responsibility to assure the validity of the test method to other low dissolved solids matrices.

1.3 The values stated in SI units are to be regarded as standard. The values given in parentheses are mathematical conversions to inch-pound units that are provided for information only and are not considered standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*³

D1066 Practice for Sampling Steam

¹ This test method is under the jurisdiction of ASTM Committee D19 on Water and is the direct responsibility of Subcommittee D19.05 on Inorganic Constituents in Water.

Current edition approved Feb. 1, 2015. Published March 2015. Originally approved in 1982. Last previous edition approved in 2008 as D4191 – 08. DOI: 10.1520/D4191-15.

² Platte, J. A., and Marcy, V. M., "A New Tool for the Water Chemist," *Industrial Water Engineering*, May 1965.

³ For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

D1129 Terminology Relating to Water

D1193 Specification for Reagent Water

D2777 Practice for Determination of Precision and Bias of Applicable Test Methods of Committee D19 on Water

D3370 Practices for Sampling Water from Closed Conduits

D3561 Test Method for Lithium, Potassium, and Sodium Ions in Brackish Water, Seawater, and Brines by Atomic Absorption Spectrophotometry

D4841 Practice for Estimation of Holding Time for Water Samples Containing Organic and Inorganic Constituents

D5810 Guide for Spiking into Aqueous Samples

D5847 Practice for Writing Quality Control Specifications for Standard Test Methods for Water Analysis

3. Terminology

3.1 *Definitions*—For definitions of terms used in this test method, refer to Terminology D1129.

4. Summary of Test Method

4.1 Sodium is determined by flame atomic absorption spectrophotometry. The sodium content is determined by aspirating the low solids sample directly with no sample pretreatment.

5. Significance and Use

5.1 Sodium salts are very soluble, and sodium leached from soil and rocks tends to remain in solution. Water with a high ratio of sodium to calcium is deleterious to soil structure.

5.2 Sodium is not particularly significant in potable water except for those persons having an abnormal sodium metabolism, but water supplies in some areas contain sufficient sodium to be a factor in the planning of sodium-free diets.

5.3 The use of sodium salts is common in industry; therefore, many industrial wastewaters contain significant quantities of sodium. For high-pressure boiler feed-water even trace amounts of sodium are of concern.

6. Interferences

6.1 In the analysis of low-solids water, interferences are usually negligible.

7. Apparatus

7.1 *Atomic Absorption Spectrophotometer* for use at 589.6 nm.

*A Summary of Changes section appears at the end of this standard

NOTE 1—The manufacturer’s instructions should be followed for all instrumental parameters. Wavelengths other than 589.6 nm may be used only if they have been determined to be equally suitable.

7.2 *Sodium Hollow-Cathode Lamps*—Multielement hollow cathode lamps are available and also have been found satisfactory.

7.3 *Pressure-Reducing Valves*—The supplies of fuel and oxidant shall be maintained at pressures somewhat higher than the operating pressure of the instrument by using suitable valves.

8. Reagents

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society where such specifications are available.⁴ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Specification **D1193**, Type I. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the bias and precision of the test method. Type II water was specified at the time of round robin testing of this test method.

8.3 *Sodium Solution, Stock* (1.0 mL = 1.0 mg Na)—Commercially purchase or dry sodium chloride to constant weight at 105°C. Dissolve 2.5418 g of the dry sodium chloride (NaCl) in water and dilute to 1 L with water.

NOTE 2—Certified sodium stock solutions of appropriate known purity are commercially available through chemical supply vendors.

8.4 *Sodium Solution, Standard* (1.00 mL = 0.1 mg Na)—Dilute 100.0 mL of sodium stock solution to 1 L with water.

8.5 *Oxidant: Air*, which has been passed through a suitable filter to remove oil, water, and other foreign substances is the usual oxidant.

8.6 *Fuel: Acetylene*—Standard, commercially available acetylene is the usual fuel. Acetone, always present in acetylene cylinders can affect analytical results. The cylinder should be replaced at 345 kPa (50 psi) to avoid acetone carryover.

8.6.1 **Warning**—“Purified” grade acetylene containing a special proprietary solvent rather than acetone should not be used with poly(vinyl chloride) tubing as weakening of the tubing walls can cause a potentially dangerous situation.

9. Sampling

9.1 Collect the samples in accordance with the applicable ASTM test method as follows: Practices **D3370** and Practice **D1066**.

⁴ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For Suggestions on the testing of reagents not listed by the American Chemical Society, see *Annual Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

TABLE 1 Determination of Bias for Sodium in Reagent Water by Atomic Absorption

Amount Added, mg/L	Amount Found, mg/L	S_t	S_o	Bias %
0.20	0.184	0.053	0.024	-8.0
1.20	1.25	0.094	0.028	+ 4.17
2.70	2.81	0.129	0.081	+ 4.07

10. Standardization

10.1 Prepare 100 mL each of a blank and at least four standard solutions to bracket the expected sodium concentration range of the samples to be analyzed by diluting the standard sodium solution (8.4) with water. Prepare the standards each time the test is to be performed or as determined by Practice **D4841**. Select the standards to give zero, middle, and maximum points for an analytical curve.

10.2 Analyze at least four working standards containing concentrations of sodium that bracket the expected sample concentration, prior to analysis of samples, to calibrate the instrument. Aspirate the blank and the standards and record the instrument readings. Aspirate water between standards.

10.3 Read directly in concentration if this capability is provided with the instrument or prepare an analytical curve by plotting the absorbance versus concentration for each standard on linear graph paper.

11. Procedure

11.1 Aspirate each sample and determine its absorbance or concentration. Aspirate water between samples.

12. Calculation

12.1 Calculate the concentration of sodium in the samples, in milligrams per litre, by either referring the absorbance obtained for each sample by reading directly in concentration if the capability is provided with the instrument, or prepared an analytical curve (10.3).

12.2 If an aliquot of diluted sample was analyzed, multiply the concentration of sodium, in milligrams per litre, by the appropriate dilution factor.

13. Precision and Bias⁵

13.1 The overall and single-operator precision of this test method for eight laboratories, which include a total of twelve operators analyzing each sample on three consecutive days, within its range for reagent water varies with the quantity being measured according to **Table 1**.

NOTE 3—Only reagent water was used to obtain the precision statement since this test method is designated for the determination of low amounts of sodium in low-solids water.

13.2 Recoveries of known amounts of sodium (from sodium chloride) in the series of prepared standards for the same laboratories and operators were as given in **Table 1**.

⁵ Supporting data have been filed at ASTM International Headquarters and may be obtained by requesting Research Report RR:D19-1080. Contact ASTM Customer Service at service@astm.org.